

9th Science Lesson 16 Notes in English

16] Applied Chemistry

Introduction

- Food, medicines, cosmetics, dress materials and gold covering ornaments are some of the items used in our day to day life.
- They may differ in nature and applications. But all these are associated with **chemistry**.
- They are made of **synthetic / natural chemicals**. We face lot of difficulties in different means to lead our day to day life.
- Such difficulties make chemists to come out of new ideas and theories.
- For example, when people suffered from diseases, new chemical compounds were synthesized and used as drugs.
- **New techniques** were also developed to **diagnose diseases**.
- When farmers suffered due to low crop yield and pest-related problems in crop field, chemists developed **new chemical fertilizers** and pesticides to combat these issues.
- Thus **chemical principles** and **theories** are applied to various fields in order to achieve specific results or to solve real-world problems. This is called **applied chemistry**.
- In this lesson, let us discuss various branches of applied chemistry and their significance.

Nanochemistry

- We know that the size and shape of materials influence their characteristics.
- Scientists found that materials having size about $1/1,000,000,000$ metre show special characteristics.
- Then they started producing such kind of materials and studied the effect of size on properties.
- Thus a new branch of chemistry called '**Nanochemistry**' was developed.
- Nanochemistry is a branch of **nanoscience**, that deals with the chemical applications of nanomaterials in nanotechnology.
- It involves synthesis and manipulation of materials at atomic and molecular level and the study of their physical and chemical properties.

Size of Nanoparticles

- The word, Nano has been derived from the Greek word '**Nanos**' which is designated to represent billionth fraction of a unit.
- For instance, **1 Nanometre = $1/1,000,000,000$ metre**.
- Can you imagine how small is a nanoparticle?
- The following examples may help to illustrate how small the **nanoscale** is.

One nanometre (nm) is 10^{-9} or 0.000,000,001 metre.

- One nanometre (nm) is 10^{-9} or 0.000,000,001 metre.
- A nanometre and a metre can be understood as the same size-difference as between golf ball and the Earth.
- Our nails grow 1 nm each second.
- The virus most usually responsible for the common cold has a diameter of 30 nm.
- A cell membrane is around 9 nm across.
- The DNA double helix is 2 nm across.
- The diameter of one hydrogen atom is around 0.2 nm.

Properties of nanomaterials

- Nanomaterials have the structural features in between those of **atoms** and **the bulk materials**.
- The properties of materials with nanometre dimensions are significantly different from those of atoms and bulk materials.
- This is mainly because the nanometre size of the materials render them, larger surface area, high surface energy, spatial confinement and reduced imperfections, which do not exist in the corresponding bulk materials.
- As the surface characteristics of nano particles are the main criteria to be considered for applications, highly sophisticated instruments like **Scanning Electron Microscope (SEM)**, **Tunneling Electron Microscope (TEM)** and **Atomic Force Microscope (AFM)** are used to analyse the surface properties of a nano particle with high resolution.

Applications of Nanochemistry

- The range of commercial products available today is very broad, including stain- resistant and wrinkle-free textiles, cosmetics, sun screens, electronics, paints and varnishes.
- **Nanochemistry** is applied in all these substances. Some of them are given below.
- The metallic nanoparticles can be used as very active **catalysts**.
- **Nano coatings** and **nanocomposites** are found useful in making variety of products such as sports equipment, bicycles and automobiles etc.
- **Nanotechnology** is applied in the production of synthetic skin and implant surgery.
- Nanomaterials that conduct electricity are being used in electronics as minute conductors to produce circuits for microchips.
- Nanomaterials have extensive applications in the preparation of cosmetics, deodorants and sun screen lotion.
- Nanoparticle substances are incorporated in fabrics to prevent the growth of bacteria.
- Nanochemistry is used in making space, defence and aeronautical devices.

Drawbacks of nanomaterials in chemistry

- Nanoparticles are unstable when they react with oxygen.
- Their exothermic combustion with oxygen can easily cause explosion.
- Because nanoparticles are highly reactive, they inherently interact with impurities as well.
- Nanomaterials are biologically harmful and toxic.
- It is difficult to synthesis, isolate and apply them.
- There are no hard-and-fast safe disposal policies for nanomaterials.

Pharmaceutical chemistry

- Pharmaceutical chemistry is the **chemistry of drugs** which utilizes the general laws of chemistry to study drugs.
- Pharmaceutical chemistry deals with preparation of drugs and the study of chemical composition, nature, behavior, structure and influence of the drug in an organism, condition of their storage and the therapeutic uses of the drugs.
- Drug discovery is the core of pharmaceutical chemistry.

Drugs

- Even though we use so many chemicals in our daily life, the chemicals used for treating diseases are termed as drug.
- The word drug is derived from the French word '**droque**' which means a dry herb.
- According to **World Health Organisation**, a drug is defined as follows:
- 'It is a substance or product that is used or intended to be used to modify or explore physiological systems or pathological states for the benefits of the recipient'.

Characteristics of drugs

- A drug must possess the following characteristics:
- It should not be toxic.
- It should not cause any side effects.
- It should not affect the receptor tissues.
- It should not affect the normal physiological activities.
- It should be effective in its action.

Sources of drugs

- The main sources of drugs are animals and plants.
- The modern manufacturers adopt many chemical strategies to synthesize drugs for specialized treatments which are more uniform than natural materials.
- The following table shows various sources of drugs.

Table 16.1 Sources of drugs

Source or Process	Drug
Plants	Morphine, Quinine
Chemical Synthesis	Aspirin, Paracetamol
Animal	Insulin, Heparin
Minerals	Liquid Paraffin
Microorganism	Penicillin
Genetic Engineering	Human growth Hormone

Types of Drugs

- Drugs fall into two general categories:
 - I The drugs that are used in the **treatment and cure** of any specific disease.
 - II The drugs that have some characteristic effect on the **animal organism**, but do not have any remedial effect for a particular disease.
- This class includes, **morphine, cocaine** etc.

1. Anaesthetics

- The **drugs** which cause loss of sensation are called **Anaesthetics**.
- They are given to patients when they undergo **surgery**.

(a) Types of Anaesthetics

- When patients undergo a major surgery in internal organs, some anaesthetics are given so that they lose sensation completely.
- But when they undergo a minor surgery in a specific part of the body, anaesthetic is given to lose sensation around that particular part.
- Based on this, there are two classes of anaesthetics as given below.

General anaesthetics:

- They are the agents, which bring about loss of all modalities of sensation, particularly pain along with '**reversible**' loss of consciousness.
- For example, when a surgery is carried out on internal organs, these anaesthetics are given.
- The patient loses consciousness for specific period of time (depending on the duration of surgery) and get it back later.

Local anaesthetics:

- They prevent the pain sensation in localised areas without affecting the degree of consciousness.

- For example, dentist give patients this kind of anaesthetics when carry out a minor surgery in teeth.

(b) Chemicals as Anaesthetics:

- There are three major chemicals which are used as anaesthetics.
- They are: **Nitrous Oxide (N₂O)**: It is a colourless, non- irritating, inorganic gas.
- It is the safest of the anaesthetic agents. This is used after mixing general anaesthetics like ether.
- **Chloroform (CHCl₃)**: It is a volatile liquid.
- It has pleasant smell and sweet taste.
- With oxygen it forms a toxic carbonyl chloride. Hence it is not used now.
- **Ether**: Diethyl ether or simple ether (C₂H₅-O-C₂H₅) is a volatile liquid.
- This is mixed with a stabilizer, 0.002% propyl halide.
- After absorption by tissues it attacks the central nervous system and makes the patient unconscious.

2. Analgesics

- Analgesics are the compounds which relieve all sorts of pains without the loss of consciousness.
- These are also called as pain killer or pain relievers.
- These are effective in healing headaches, myalgia and arthralgia.
- **Aspirin** and **Novalgine** are the commonly used analgesics.
- Aspirin acts both as antipyretic as well as analgesic.
- Certain **narcotics** (which produce sleep and unconsciousness) are also used as analgesics.
- The analgesics are given either orally or applied externally. In general, externally applicable pain killers come as 'gels'.

3. Antipyretics

- Antipyretics are the compounds which are used for the purpose of reducing fever (lowering the body temperature to the normal).
- They are taken orally as **tablets** and **capsules**.
- The most common antipyretics are, aspirin, antipyrine, phenacetin, and paracetamol.

4. Antiseptics

- Antiseptic is a substance that prevents infections caused by disease causing microorganisms or pathogens.
- **Antiseptics** either kill the micro-organism or prevent their growth.
- Antiseptics are used externally to cleanse wounds and internally to treat infections of the intestine and bladder.

- **Iodoform** (CHI₃) is used as an antiseptic and its 1% solution is a disinfectant.
- 0.2 % solution of phenol acts as an antiseptic and its 1% solution is a disinfectant.
- **Hydrogen peroxide** is a minor antiseptic mainly used for cleansing wounds.

5. Antimalarial

- Malaria is a **vector borne disease** which causes shivering and fever. It raises the body temperature to 103-106 °F.
- It causes physical weakness with the side-effects in liver and also causes anaemia.
- Extracts of roots and stems of certain plants are extensively used as antimalarial.
- **Quinine** is a natural antimalarial obtained from Cinchona bark.



Figure 16.1 Cinchona Bark

- The last antimalarial discovered in 1961 is pyrimethamine.
- However, **quinine**, **primaquine** and **chloroquine** are some of the best antimalarials.
- Chloroquine is used specially to control malarial parasites such as plasmodium ovale, plasmodium vivax etc.
- It is not used in curing the disease. It is used as an additive with other antimalarial drugs.

6. Antibiotics

- Many microorganisms (bacteria, fungi and molds) produce certain chemicals which inhibit the growth or metabolism of some other disease causing micro organisms.
- Such chemical compounds are known as **antibiotics**.
- These need to be present only in low concentration to be effective in their antibiotic action.
- The first antibiotic '**penicillin**' was discovered by **Alexander Fleming** in 1929, from the mold Penicillium notatum.
- Penicillin is extensively used for rheumatic fever, narrowing of heart wall, bronchitis, and pneumonia etc.
- There are three main sources of antibiotics:

I Bacteria

II Fungi and

III Actinomycetes.

- The original antibiotics, like a lot of today's antibiotics, are derived from natural sources.
- Certain plant extracts, essential oils, and even foods have antibiotic properties. Example: Honey, garlic, ginger, clove, neem and turmeric.

7. Antacids

- Quite often, after eating oily and spicy food, one may feel uncomfortable due to some burning sensation in stomach / food pipe.
- This is due to imbalance in the acidity in the stomach.
- Certain drug formulations provide relief from such burning sensation.
- These are known as **antacids**.
- Antacids are available in tablet as well as gel / syrup forms.
- These antacids contain magnesium and aluminium hydroxides, in addition to flavouring agents and colour.

Electrochemistry

- We use so many electronic devices like mobile phone, and electrical devices like torch light, in our daily life.
- Electricity produced by the battery is the key factor which makes these devices to function.
- But how does battery produce electricity?
- Because it contains some chemicals in it.
- The chemical reactions (chemical energy) that take place in the battery produce electricity (electrical energy).
- So, when scientists realized that **chemical energy can be converted into electrical energy** and vice versa, another branch of applied chemistry was developed. It is **Electrochemistry**.
- Electrochemistry is a branch of chemistry which deals with the relation between electrical energy and chemical change.
- It is mainly concerned with the processes taking place between the electrode and solution having ions called electrolyte.

Electrochemical Cell

- So many chemical reactions take place around us.
- Do all they produce electricity? No.
- Only redox reactions that take place in a specific device can produce electricity.
- The device that make use of a **chemical change to produce electricity or electricity to produce chemical change** is called **Electrochemical Cell**.

(a) Components of Electrochemical Cell

- An electrochemical cell may comprise of the following two major components.
- **Electrode:** It is a solid electrical conductor made of metal (sometimes non-metal like graphite).
- A cell consists of two electrodes. One is called **Anode** and the other is called **Cathode**.
- **Electrolyte:** It is made up of solutions of ions or molten salts which can conduct electricity.

(b) Cell reactions

- An electrochemical cell involves two reactions simultaneously.

Oxidation: As we know already, an oxidation is **loss of electron**. In electrochemical cells, oxidation takes place at anode.



Reduction: It involves gain of electron. Reduction takes place at cathode



- **Oxidation:** As we know already, an oxidation is loss of electron. In electrochemical cells, oxidation takes place at anode.
- **Reduction:** It involves gain of electron. Reduction takes place at cathode
- Since both the reactions take place simultaneously, the inter conversion of electrical and chemical energy in electrochemical cells involves a **redox reaction**.

(c) Types of Electrochemical Cell

- Based on the nature of the energy conversion, electrochemical cells are broadly classified as below.

Galvanic Cell

- It is an electrochemical cell which converts chemical energy into electrical energy i.e. it **produces electricity from chemical reactions**.
- It consists of two half cells namely **anodic half-cell** and **cathodic half-cell**.
- In anodic half-cell, the anode is in contact with its electrolyte whereas in cathodic half-cell, the cathode is in contact with its electrolyte.
- The anode and cathode are connected by a conductor wire. The electrolytes of half-cells are connected through a tube containing a saturated salt solution.
- It is called **salt bridge**. Thus in galvanic cell, both the half-cells are kept separately but stay connected electrically.
- How does a galvanic cell produce electricity?
- At anode, oxidation takes place which releases electrons.

- These electrons are attracted by cathode and hence the electrons flowing from anode to cathode are gained in reduction reaction.
- As long as the redox reaction proceeds, there is a flow of electrons and hence electricity.

Electrolytic Cell

- It is an electrochemical cell which **converts electrical energy into chemical energy** i.e. in electrolytic cells, electricity is used to bring about chemical reactions.
- Here, both anode and cathode are in contact with same electrolyte and thus the half-cells are not separated.
- As seen in **galvanic cells**, electrolytic cell also involves **redox reaction**.
- We get electricity from galvanic cells. But electrolytic cells use electricity. Then how are they useful?

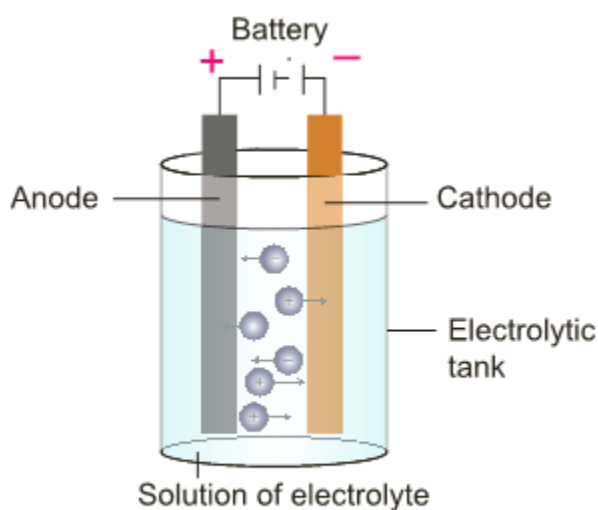


Figure 16.2 Electrolytic Cell

- In electrolytic cells, when electricity is passed to the electrolyte, it dissociates into its constituent ions.
- These ions undergo redox reaction forming the respective elements. This phenomenon is called **Electrolysis**.
- So, electrolysis is a process by which an electrolyte is decomposed into its constituent elements by passing electricity through its aqueous solution or fused (molten) state.

(b) Significance of electrochemistry

- The subject of electro chemistry is of great significance. Some of its applications are given below.
 - I. It has been used to discover important technical processes for the production and purification of non-ferrous metals, and for the electro- synthesis of organic compounds.
 - II. Electrochemistry is used to predict whether a particular reaction will occur or not.

- III. The detection of alcohol in drunken drivers is possible through the electrochemical redox reaction of ethanol.
- IV. Production of metals like aluminum and titanium from their ores involve electrochemical reactions.
- V. Lead acid batteries, lithium-ion batteries and fuel cells are based on electrochemical cells.
- **Fuel cell** is used to bring about direct conversion of chemical energy into electrical energy.

Radiochemistry

- You have studied in previous chapters that elements can exist in nature as their isotopes.
- **Isotopes** are atoms with same number of protons and electrons and a different number of neutrons.
- Some isotopes are stable and stay forever. These are the elements that we see around us and find in nature.
- However, some isotopes are unstable and they undergo disintegration by losing their energy in the form of radiation.
- As we studied earlier, every element tries to attain stability by sharing, losing or gaining electrons (octet rule).
- Thus, the unstable isotopes of elements lose their energy in the form of radiation to become stable.
- This phenomenon is called **radioactive decay**.
- The isotope which undergoes radioactive decay is called **radioactive isotope or radioisotope**.
- This property of isotopes is known as **radioactivity**.

Table 16.2 Radioisotope in Diagnosis

Radioisotope	Diagnosis used for
Iodine-131	Location and detection of brain tumor, thyroid gland disorder.
Sodium-24	Location of blood clot and circulation disorders, pumping action of heart.
Iron-59	Diagnosis of anaemia, pregnancy disorder.
Cobalt-60	Diagnosis of cancer.
Hydrogen-3	Water content of the human body.

- **Radiochemistry** is the study of chemistry of radioactive and non-radioactive isotopes. It includes both natural and artificial isotopes.

- Radiochemistry mainly deals with application of radioisotopes to study the nature of chemical reactions of non-radioactive isotopes of elements and applications of radioisotopes to various fields.

Table 16.3 Radioisotope in Treatment

Radioisotope	Treatment used for
Gold-198	Cancer
Iodine-131	Hyperthyroidism and cancer
Phosphorous-32	Blood disorder and skin disease
Cobalt-60	Cancer

Applications of Radiochemistry

- Radioisotopes can easily be detected and estimated quantitatively. So they are used in radiochemistry for various applications.
- Radiochemistry mainly deals with study of chemical reactions of non-radioactive isotopes using radioisotopes.
- In addition to that it could find applications in medical field and environmental management also.
- Let us list important applications of radioisotopes.
- **Radiocarbon dating:** It is a method by which the age of fossil wood or animal is determined using C-14 isotope.
- **Diagnosis:** Radioisotopes are found very useful to diagnose and understand many diseases.
- **Radiotherapy:** Radioactive isotopes are used in the treatment of many diseases. This kind of treatment is called radiotherapy.

Dye Chemistry

- Human is always fascinated by colours, because we are living in a colourful world. We could see so many colours in plants and their flowers.
- We eat coloured food stuffs and use numerous coloured materials in our daily life. Do you know how do they get coloured?
- Because they contain some kind of chemicals in them which are called **colourants**.
- The uses of colourants by mankind for painting and dyeing dates back to the dawn of civilization.
- Until the middle of the 19th century, all colourants applied were from natural origin.
- For example, inorganic pigments such as soot, manganese oxide, hematite were used as colourants.
- **Organic natural colourants** have also a timeless history of application, especially for colouring textiles.

- The organic compounds that are used as colourants are called **dyes**.
- These dyes are all aromatic compounds, originating from plants and also from insects, fungi and lichens.
- After the evolution of modern organic chemistry, many kinds of synthetic dyes were prepared and used by mankind.
- **Dye chemistry** is the study of such kind of dyes.
- It provides us information on theory, structure, synthesis and applications of synthetic dyes.

Characteristics of Dyes

- All coloured compounds are not dyes.
- Dyes are those coloured compounds which can be firmly fixed in fabrics by chemical or physical bonding.
- So, a dye should have the following characteristics:
- It should have a suitable colour.
- It should be able to fix itself or be capable of being fixed to the fabric.
- It should be fast to light.
- It should be resistant to the action of water, dilute acids and alkalies.

Classification of dyes

- Nowadays, practically all the dyes are synthetic, and are prepared from aromatic compounds obtained from coal tar.
- Therefore, such dyes are sometimes called as **coal tar dyes**.
- But, they may differ in their basic structure and the way of application.
- So, dyes are classified in two ways: one, based on the method of application and other based on their parent structure.

(a) Based on method of application

- **Acid dyes:** These are acidic in nature and used for dyeing animal fibres and synthetic fibres. These can be used for protein fibres such as wool and silk. E.g. Picric acid, Naphthol yellow-s
- **Basic dyes:** These are basic dyes containing basic group (-NH₂, -NHR, -NR₂). They are used for dyeing animal fibres and plant fibres.
- **Mordant dyes or Indirect dyes:** These dyes have a poor affinity for cotton fabrics and hence do not dye directly.
- They require pretreatment of the fibre with a mordant.
- **Mordant** (latin : mordere = to bite) is a substance which can be fixed to the fibre and then can be combined with the dye to form an insoluble complex called **lake**.
- Aluminium, chromium, and iron salts are widely used as mordants. E.g. alizarin.
- **Direct dyes:** They have high affinity for cotton, rayon and other cellulose fibre. So, they are applied directly as they fix firmly on the fabric. E.g. Congo red
- **Vat dyes:** It can be used only on cotton and, not on silk and wool.



Figure 16.3 Vat dyes

- This dyeing is a continuous process and is carried out in a large vessel called **vat**.
- So, it is called as vat dye. E.g. Indigo

(b) Based on Structure

- Based on the structure, dyes are classified as:
- Azo dyes
- Diphenyl methane dye
- Triphenyl methane dye
- Phthalein dye
- Anthraquinone dye
- Indigo dyes
- Phthalo cyanine dye
- Nitro and nitroso dyes

Agricultural and Food Chemistry

Agricultural Chemistry

- Agricultural chemistry involves the application of chemical and biochemical knowledge to agricultural production, processing of raw materials into foods and beverages, and environmental monitoring and remediation.
- It deals with scientific relation between plants, animals, bacteria and environment

(a) Goals of agricultural chemistry

- Indian chemists and biochemists applied their knowledge and developed modernized agricultural practices which involve use of synthetic fertilizers, genetically modified crops, and equipments.
- It aims at producing sufficient nutritious food and feed the population in a sustainable way while being responsible stewards of our environment and ecosystem.

- Based on the issues and challenges in agricultural production, agricultural chemistry mainly focusses to achieve the following:
 - I. **Increase in crop yield and livestock**
 - II. **Improvement of food quality**
 - III. **Reducing cost of food production**

(b) Applications of Agricultural Chemistry

- Chemical principles and reactions are most widely used in agriculture in order to increase yield, to protect crops from diseases and to simplify the practice of agriculture. Various applications are give below.
- **Soil Testing:** Crop lands may have different kinds of soil with varying pH.
- **Soil pH** is one of the main criteria to be considered for the selection of crop or remediation of soil.
- **Soil testing** involves determination of pH, porosity and texture.
- **Chemical Fertilizers:** Fertilizers are chemical compounds added to crop field for supplying essential micro and macro nutrients required for crop growth.
- Ammonium nitrate, calcium phosphate, urea, NPK (Nitrogen, Phosphorous and Potassium), etc. are some of the fertilizers.
- Depending on the nature of soil, these fertilizers are used singly or as mixtures.
- **Pesticides and Insecticides:** Crops are prone to diseases caused by pests and insects.
- Chemically synthesized pesticides and insecticides are used to solve these issues.
- Chlorinated hydrocarbons, organophosphates and carbamates are used as pesticides and insecticides.

Food Chemistry

- Food is one of the basic needs of human and animal.
- The food we eat also are made of chemicals. Any human might require the following three kinds of food:
 - **Body building foods:** These are required for physical growth of body. E.g. Proteins
 - **Energy giving foods:** These are the foods that supply energy for the functioning of parts human body. E.g. Carbohydrates and Fats.
 - **Protective foods:** These protect us from deficiency diseases. E.g. Vitamins and Minerals
- Every human requires all these three kind of foods in right proportion for the smooth functioning of the body.
- The diet that contain all these three foods in right proportion is called **Balanced diet**.
- **Food chemistry** is the chemistry of foods which involves the analysis, processing, packaging, and utilization of materials including bioenergy for food safety and quality.

Table 16.4 Food additives

Type of additive	Function of the additive	Example
Preservatives	They protect food from spoilage by microorganism in storage.	Vinegar, Sodium benzoate, benzoic acid, sodium nitrite
Colourants	They give pleasant colours to food	Carotenoids, Anthocyanin, Curcumin
Artificial Sweeteners	They add sweet taste to food	Saccharin, Cyclamate
Flavor enhancers	They are used to enhance the flavour of food items	Monosodium glutamate, Calcium diglutamate
Antioxidants	They prevent the oxidation of food. They protect us against cardiovascular disease.	Vitamin C, Vitamin E, Carotene

(a) Goals of food chemistry

- The main goal of food chemistry is to cater **the needs of quality food** to the population in a sustainable way.
- In basic research, **food chemists** study the properties of proteins, fats, starches, and carbohydrates, as well as micro components such as additives and flavourants, to determine how each works in a food system.
- In application research, they often develop new ways to use ingredients or new ingredients altogether, such as fat or sugar replacements.

(b) Chemicals in Food

- Food we eat in our day to day life contains natural or synthetic chemicals. They serve different functions in human body.
- **Nutrients:** They are the most essential chemicals present in food.
- They are required for the growth, physiological and metabolic activities of body.
- They are natural or synthetic. E.g. Carbohydrates, proteins, vitamins and minerals.
- **Food additives:** These are the chemicals added to food for specialized functions.
- The various types of additives of foods are given in Table 16.4.

Forensic Chemistry



Figure 16.4 Crime detection

- Forensic chemistry applies scientific principles, techniques, and methods to the investigation of crime.
- Our daily newspaper is carrying a lot of news on incidents of criminal activities such as robbery, murder, sexual harassment, etc.
- How the crime department investigate and analyse it?
- In real life the collection and analysis of evidence involve painstaking care and rigorous application of scientific principles.

Forensic Chemists in Criminal Investigation

- In general, forensic chemists work in four steps in the investigation of crime.
- **Collection of Evidences:** They collect physical evidences such as knife, instruments, materials, etc in a systematic way and uncover their information using chemistry.
- **Analysis of evidences:** In criminal cases, chemists analyze substances such as blood and DNA to attempt to determine when and by whom the crime was committed.

Collaboration:

- To solve the crime, they discuss with other fellow investigators like police officers, detective and other forensic scientists.
- **Report of findings:** Finally, they prepare a report of the conclusion of the analysis.

Method of Forensic Chemistry

- The world of forensic chemistry, focusing on the theory and processes of chemistry in forensic analysis shows the role that chemistry plays in criminal investigations.
- The following are some methods used in crime investigation by a forensic chemistry lab.
- **Finger print:** Finger print is one of the most important evidences in crime investigation.
- Fingerprints on smooth surfaces can often be made visible by the application of light or dark powder, but fingerprints on cheque or other documents are often occult (hidden).
- Occult fingerprints are sometimes made visible by the use of ninhydrin, which turns purple due to reaction with amino acids present in perspiration.

- Fingerprints or other marks are also sometimes made visible by exposure to high-powered laser light.
- **Cyanoacrylate** ester fumes from glue are used with fluorescent dyes to make the fingerprints visible.



Figure 16.5 Finger print

- **Biometrics:** The science that involves the study and analysis of human body prints is known as biometrics.
- The biometric system compares the body prints to the specimen data stored in the system to verify the identity of a person.
- **Alcohol test:** Drinkers can be easily identified by the use of applied chemistry.



Figure 16.6 Alcohol test

- The person being tested blows through a tube, which bubbles the breath through a solution of chemicals containing sulfuric acid, potassium dichromate, water, and silver nitrate.
- Oxidation of the alcohol results in the reduction of dichromate to chromic ion, with a corresponding change in colour from orange to green.
- An electrical device employing a photocell compares the colour of the test solution with a standard solution, giving a quantitative determination of the alcohol content.
- The test provides a quick and reproducible determination of the amount of alcohol in a person's breath and is a numerical measure of the amount of alcohol in the bloodstream.

Applications of Applied Chemistry

- Many of the advantages of applied chemistry are around us.
- It is inevitable.
- Applied chemistry has given us innumerable synthetic materials to lead our day to day life.
- The applied chemistry makes a most important contribution to our society.
- It makes a major contribution to the country's economic development, and plays vital role worldwide.
- The products of applied chemistry are so widespread that they are used in our daily.

MORE TO KNOW:

Uranium

- Uranium in the ground can decay into radon gas which can be very dangerous to humans.
- It is thought to be the second leading isotope to cause lung cancer.

Agriculture

- 40% of today's global population works in agriculture, making it the single largest employment in the world.

EXTRA POINTS:

- **Anaesthetics:** Drugs which cause loss of sensation.
- **Antipyretics:** Compounds which are used for the purpose of reducing fever.
- **Antiseptic:** Substance that prevents infections caused by disease causing microorganisms or pathogens.
- **Antibiotics:** Chemical compounds which was produced by many microorganisms (bacteria, fungi and moulds) which inhibit the growth or metabolism of some other disease causing microorganism.
- **Antacids:** These are certain drug formulations which provide relief from burning sensation.
- **Drug:** The chemicals used for treating diseases
- **Electrochemical Cell:** The device that make use of a chemical change to produce electricity or electricity to produce chemical change.
- **Electrolyte:** Solutions of ions or molten salts which can conduct electricity
- **Pharmaceutical Chemistry:** It is the study of drugs and it involves drug development.